Zuschriften

Magnetic Nanowires

DOI: 10.1002/ange.200503486

Magnetic-Dipolar-Interaction-Induced Self-Assembly Affords Wires of Hollow Nanocrystals of Cobalt Selenide**

Jinhao Gao, Bei Zhang, Xixiang Zhang, and Bing Xu*

Self-assembly, an attractive and practical methodology, allows the formation of a wide range of nanostructures for promising applications—for example, nanoparticle arrays for new opti-

[*] J. Gao, Prof. B. Xu

Department of Chemistry

The Hong Kong University of Science & Technology

Clear Water Bay, Hong Kong (China)

Fax: (+852) 2358-1594 E-mail: chbingxu@ust.hk

B. Zhang, Prof. X. Zhang Department of Physics

The Hong Kong University of Science & Technology

Clear Water Bay, Hong Kong (China)

[**] This work was partially supported by RGC (Hong Kong) and EHIA



Supporting information for this article is available on the WWW under http://www.angewandte.org or from the author.

cal band-gap materials^[1] or high-density magnetic recording media, [2] self-assembled monolayers (SAM) for nanometerthick films on a variety of substrates, [3] and nanofibers of selforganized small molecules or oligopeptides for drug delivery and tissue engineering.^[4] Hydrogen bonding and/or van der Waals interactions are usually the driving forces that result in self-assembly at the nanoscale. Despite extensive investigation, it is still difficult to control and predict the nanostructures resulting from these two types of forces. Although several recent reports have shown that magnetic dipolar interactions induce rather predictable nanoassemblies, [6,7] methods for maintaining these assemblies when the magnetic force decreases or vanishes have been less thoroughly investigated. Therefore, we decided to use the magnetic dipoles inherently associated with magnetic nanoparticles^[8] to form 1D assemblies of hollow nanocrystals of semiconductors.[9]

We chose to make wires of hollow nanocrystals of CoSe₂ to demonstrate the concept, because 1) the isolation of a nanonecklace of cobalt nanocrystals on a substrate from a dispersion after removal of the solvent^[7] indicated that magnetic dipolar interactions are sufficient to maintain 1D assemblies of nanocrystals in solution; 2) the recently reported formation of hollow CoSe nanocrystals from cobalt nanocrystals through the Kirkendall effect^[10] suggested that a similar process could be used to generate hollow CoSe₂ nanocrystals from 1D assemblies of cobalt nanocrystals; 3) hollow nanostructures have received considerable attention because they exhibit properties that differ from their solid counterparts; [11] and 4) nanowires of hollow nanocrystals of CoSe₂ are unknown. Herein, we show that when their sizes reach approximately 20 nm cobalt nanocrystals self-assemble into wires in solution at temperatures as high as 455 K. Through the nanoscale Kirkendall effect, [10,12] wires of hollow CoSe₂ nanocrystals form without loss of the preassembled nanostructure induced by the magnetic dipolar interactions of the cobalt nanocrystals. In control experiments, when the size of the cobalt nanocrystals is 6 nm, or when 20-nm cobalt nanocrystals are in an alternating magnetic field, no wires of hollow CoSe₂ nanocrystals form. These observations confirm that the magnetic interaction plays a key role in the formation of the nanowires of hollow CoSe₂ crystals. Several types of ferromagnetic metallic nanocrystals (iron, cobalt, nickel, FePt, or CoPt) can be prepared readily, and their selfassembly depends on the magnetic dipolar interactions of particles of a certain size. Therefore, the method shown herein should be a general route to 1D assemblies of hollow

The synthesis of wires of hollow CoSe₂ nanocrystals is easy and straightforward. Cobalt nanocrystals of approximately 20 nm in diameter were prepared according to a modified literature procedure^[13] and then dispersed in a solvent (for example, 1,2-dichlorobenzene) by using trioctylphosphine oxide (TOPO) as a surfactant. The injection of a dichlorobenzene solution of selenium into the dispersion of cobalt nanocrystals under reflux at 455 K resulted in a black dispersion of CoSe₂ nanocrystals after approximately 30 min. Transmission electron microscope (TEM) images reveal that the CoSe₂ nanocrystals form wires of interconnected hollow spheres (Figure 1a). High-resolution TEM images (inset, Figure 1a) and analysis of the electron diffraction pattern of the sample suggest that these hollow spheres are crystalline. Scanning electron microscope (SEM)

Figure 1. a) TEM image (inset: HR-TEM image), and b) SEM image of the wires of hollow CoSe₂ nanocrystals.

images indicate that the wires of hollow CoSe2 nanocrystals are the major nanostructure formed (Figure 1b). This observation suggests that the wires form in high yield, which may allow large-scale production. The X-ray diffraction (XRD) pattern of the CoSe₂ nanocrystals (Figure 2) matches that of

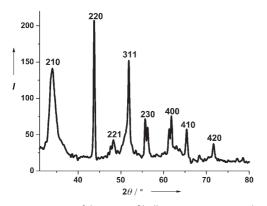
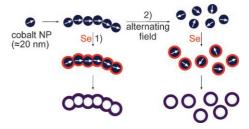


Figure 2. XRD pattern of the wires of hollow CoSe₂ nanocrystals.

the primitive cubic phase of CoSe₂ (Powder Diffraction File #090234). The X-ray photoelectron spectrum (XPS) of the CoSe₂ nanocrystals shows peaks at 778.7 and 54.4 eV, corresponding to the binding energies of the cobalt 2p_{3/2} and selenium 3d_{5/2} states.^[14] In addition, X-ray fluorescence (XRF) analysis indicates that the ratio of Co:Se in the sample is approximately 1:1.80, which suggests that other cobalt selenide phases (for example, CoSe and Co₃Se₄) may be present in small amounts.^[15]

Additional TEM analysis of the dispersion of cobalt nanocrystals prior to the addition of selenium showed that the cobalt nanocrystals assemble into wires (see Supporting Information), which are similar to the assemblies of hollow CoSe₂ nanocrystals. We also observed a small amount of necklace structures in both the samples of cobalt nanocrystals and hollow CoSe₂ nanocrystals (see Supporting Information). Furthermore, the sizes and alignment of the holes of the hollow CoSe₂ nanocrystals match with the sizes and alignment of the cobalt nanocrystals in the 1D assembly. These results lead to the proposed mechanism for the formation of wires of hollow CoSe₂ nanocrystals (Scheme 1). Each cobalt nano-



Scheme 1. Formation of hollow CoSe2 nanocrystals from cobalt nanoparticles (NP) in the 1) absence and 2) presence of an alternating magnetic field.

crystal behaves as a single magnetic dipole, and the cobalt nanocrystals form chains or necklaces in solution because of strong magnetic dipolar interactions. After being added into the dispersion, selenium reacts with cobalt, and a thin layer of selenium or CoSe₂ grows around the wires of cobalt nanocrystals, without disintegration of the wires. Further reaction between selenium and cobalt through the nanoscale Kirkendall process maintains the preassembled nanostructures and affords a 1D assembly of hollow CoSe2 nanocrystals. According to this mechanism, disruption (process 2) in Scheme 1) or reduction of the magnetic dipolar interactions before the addition of selenium should prevent the formation of the wires of hollow CoSe₂ nanocrystals. To illustrate this point, we conducted the two following control experiments.

First, we chose cobalt nanocrystals with a smaller size (6 nm) as starting materials, because they have weaker magnetic dipolar interactions than cobalt nanocrystals with diameters of 20 nm. Under the conditions described above, the reaction between 6-nm cobalt nanocrystals and selenium only affords monodispersed hollow nanocrystals (Figure 3a). No 1D assemblies of nanocrystals are observed, probably because the magnetic dipolar interactions between these small cobalt nanocrystals are too weak to form chain or necklace nanostructures at 455 K. Second, we synthesized hollow CoSe₂ nanocrystals in an alternating magnetic field,

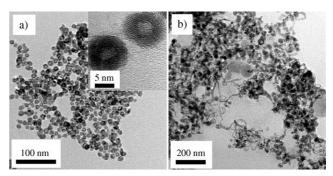


Figure 3. a) TEM image (inset: HR-TEM image) of the monodispersed hollow CoSe₂ nanospheres synthesized from 6-nm cobalt nanocrystals; b) TEM image of the hollow CoSe₂ nanostructures synthesized in an alternating magnetic field.

1243

Zuschriften

 $B(t) = B_0 \sin(\omega t)$, where $B_0 = 35$ G and $\omega = 2\pi f$ with f = 50 Hz, generated by a Helmholtz coil. In this field, the reaction of 20nm cobalt nanocrystals with selenium gives only a small number of individual hollow nanocrystals of CoSe2, in addition to a small number of selenium nanowires (Figure 3b).[16] This result is consistent with process 2) proposed in Scheme 1. Because of the head-to-tail arrangement of the magnetic dipoles of the nanoparticles, it is difficult to "switch" all of the dipoles at the same time. The breaking of the chain is due to the vibrational magnetic torque applied, which disrupts the dipolar interactions between the particles. Once the chain is broken, the alternating magnetic torque causes the particles to rotate and prevents them from reconnecting. This mechanism agrees with our observations. These two experiments confirm that magnetic dipolar interactions play a key role in the formation of 1D assemblies of hollow CoSe₂ nanocrystals. The use of magnetic dipolar interactions, which can be easily modified through the application of an alternating magnetic field, in the formation of nanostructures provides a useful alternative to the use of electronic dipolar interactions (as applied in the assembly of nanowires of CdSe^[5] and PbTe^[17]).

To establish the generality of this procedure, we replaced selenium by sulfur or tellurium in the reaction with 20-nm cobalt nanocrystals. As shown in Figure 4, we obtained wires

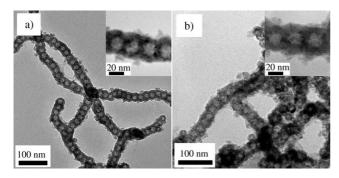


Figure 4. TEM images (insets: HR-TEM images) of the wires of hollow nanocrystals of a) Co₃S₄ and b) CoTe.

of hollow Co₃S₄ and CoTe nanocrystals, which are similar in morphology to the hollow CoSe₂ nanocrystals. The XPS of the Co₃S₄ wires shows peaks at 778.9 and 162.2 eV, corresponding to the binding energies of the cobalt 2p_{3/2} and sulfur 2p states. XRF analysis indicates that the ratio of Co:S in the sample is 1:1.32, nearly 3:4, which agrees with the ratios of the corresponding precursors. For the CoTe sample, XRF analysis indicates that the ratio of Co:Te is approximately 1:1 (see Supporting Information).

Magnetization measurements indicate that the wires of hollow CoSe₂ nanocrystals are ferromagnetic. As shown in Figure 5, the field-dependent magnetization (M(H)) of the wires of hollow CoSe₂ nanocrystals at 50 K displays hysteresis typical of a ferromagnet. The coercivity (H_c) of the wires of hollow CoSe₂ nanocrystals is approximately 600 Oe, roughly half that of the 20-nm cobalt nanocrystals (approximately 1100 Oe; see Supporting Information). Standard zero-field cooling (ZFC) and field cooling (FC) temperature-dependent magnetization (M(T)) measurements indicate bulk ferromag-

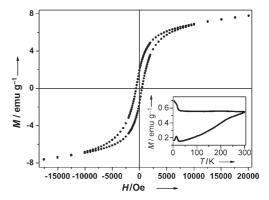


Figure 5. The hysteresis loop for the wires of hollow CoSe₂ nanocrystals at 50 K (inset: temperature-dependence of the ZFC (lower) and FC (upper) magnetization of the wires of hollow CoSe₂ nanocrystals in a magnetic field of 100 Oe).

netism. Similar M(T) curves were observed for 20-nm cobalt nanocrystals (see Supporting Information). The magnetic behavior of the wires of hollow CoSe2 nanocrystals suggests that these nanocrystals may actually form a polycrystalline wire with holes. The small peak at 14 K in the ZFC curve (Figure 5, inset) could be due to the unblocking of superparamagnetic CoSe₂ nanocrystals, which is evidenced by the FC data from 5 to 30 K. To exclude the possibility that the ferromagnetism solely originates from residual cobalt nanocrystals, we measured the magnetization of the wires of hollow CoSe₂ nanocrystals in a 1000-Oe field from 5 to 674 K (Figure 6). We used a SQUID magnetometer to collect the

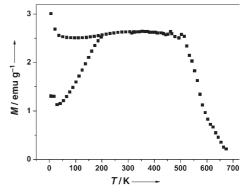


Figure 6. Temperature-dependence of the ZFC (lower) and FC (upper) magnetization of the wires of hollow CoSe2 nanocrystals in a magnetic field of 1000 Oe.

data from 5 to 300 K and a vibrating sample magnetometer to collect the data from 300 to 674 K. The Curie temperature was determined by the position of the maximum of |dM/dT|to be 570 K, which is roughly 1000 K lower than the Curie temperature of cobalt. This result suggests that the ferromagnetism of the wires may arise from CoSe2, although the detailed mechanism has yet to be established.

In summary, we have demonstrated that the nanoscale Kirkendall effect can be applied to a magnetically preassembled nanostructure to form 1D assemblies of hollow nanocrystals of cobalt chalcogenides at elevated temperatures.

1244

This procedure provides an alternative route to nanowires with a unique morphology (interconnected hollow nanocrystals) that is difficult to achieve using other means. This "magnetic-guiding" strategy is also compatible with other methodologies for creating nanostructures.^[9,18] In addition, the magnetic nanostructures produced may have potential applications in the fabrication of nanodevice components. We are currently investigating this possibility.

Received: October 3, 2005 Revised: November 21, 2005 Published online: January 18, 2006

Keywords: chalcogenides · cobalt · magnetic properties · nanostructures · self-assembly

- a) S. H. Park, D. Qin, Y. Xia, Adv. Mater. 1998, 10, 1028; b) B. Gates, Y. Xia, Appl. Phys. Lett. 2001, 78, 3178.
- [2] a) S. Sun, C. B. Murray, D. Weller, L. Folks, A. Moser, *Science* 2000, 287, 1989; b) V. F. Puntes, P. Gorostiza, D. M. Aruguete, N. G. Bastus, A. P. Alivisatos, *Nat. Mater.* 2004, 3, 263; c) T. Hyeon, *Chem. Commun.* 2003, 927.
- [3] J. C. Love, L. A. Estroff, J. K. Kriebel, R. G. Nuzzo, G. M. Whitesides, *Chem. Rev.* 2005, 105, 1103.
- [4] a) T. C. Holmes, S. de Lacalle, X. Su, G. S. Liu, A. Rich, S. G. Zhang, *Proc. Natl. Acad. Sci. USA* 2000, 97, 6728; b) G. A. Silva, C. Czeisler, K. L. Niece, E. Beniash, D. A. Harrington, J. A. Kessler, S. I. Stupp, *Science* 2004, 303, 1352; c) B. G. Xing, C. W. Yu, K. H. Chow, P. L. Ho, D. G. Fu, B. Xu, *J. Am. Chem. Soc.* 2002, 124, 14846; d) Z. M. Yang, H. W. Gu, Y. Zhang, L. Wang, B. Xu, *Chem. Commun.* 2004, 208.
- [5] a) Z. Y. Tang, N. A. Kotov, M. Giersig, Science 2002, 297, 237;
 b) Z. Y. Tang, N. A. Kotov, Adv. Mater. 2005, 17, 951.
- [6] a) J. C. Love, A. R. Urbach, M. G. Prentiss, G. M. Whitesides, J. Am. Chem. Soc. 2003, 125, 12696; b) A. R. Urbach, J. C. Love, M. G. Prentiss, G. M. Whitesides, J. Am. Chem. Soc. 2003, 125, 12704.
- [7] a) S. L. Tripp, S. V. Pusztay, A. E. Ribbe, A. Wei, J. Am. Chem. Soc. 2002, 124, 7914; b) S. L. Tripp, R. E. Dunin-Borkowski, A. Wei, Angew. Chem. 2003, 115, 5749; Angew. Chem. Int. Ed. 2003, 42, 5591.
- [8] a) T. Hyeon, S. S. Lee, J. Park, Y. Chung, H. Bin Na, J. Am. Chem. Soc. 2001, 123, 12798; b) J. Park, K. J. An, Y. S. Hwang, J. G. Park, H. J. Noh, J. Y. Kim, J. H. Park, N. M. Hwang, T. Hyeon, Nat. Mater. 2004, 3, 891; c) J. Park, E. Lee, N. M. Hwang, M. S. Kang, S. C. Kim, Y. Hwang, J. G. Park, H. J. Noh, J. Y. Kini, J. H. Park, T. Hyeon, Angew. Chem. 2005, 117, 2932; Angew. Chem. Int. Ed. 2005, 44, 2872.
- [9] Y. D. Yin, A. P. Alivisatos, Nature 2005, 437, 664.
- [10] Y. D. Yin, R. M. Rioux, C. K. Erdonmez, S. Hughes, G. A. Somorjai, A. P. Alivisatos, *Science* 2004, 304, 711.
- [11] a) J. Chen, F. Saeki, B. J. Wiley, H. Cang, M. J. Cobb, Z. Y. Li, L. Au, H. Zhang, M. B. Kimmey, X. D. Li, Y. Xia, Nano Lett. 2005, 5, 473; b) Y. Sun, Y. Xia, J. Am. Chem. Soc. 2004, 126, 3892; c) L. Rapoport, N. Fleischer, R. Tenne, Adv. Mater. 2003, 15, 651; d) L. Rapoport, Y. Bilik, Y. Feldman, M. Homyonfer, S. R. Cohen, R. Tenne, Nature 1997, 387, 791; e) N. A. Dhas, K. S. Suslick, J. Am. Chem. Soc. 2005, 127, 2368; f) D. Goll, A. E. Berkowitz, H. N. Bertram, Phys. Rev. B 2004, 70; g) R. Tenne, Prog. Inorg. Chem. 2001, 50, 269; h) Y. G. Sun, B. Mayers, Y. N. Xia, Adv. Mater. 2003, 15, 641; i) X. M. Sun, Y. D. Li, Angew. Chem. 2004, 116, 3915; Angew. Chem. Int. Ed. 2004, 43, 3827; j) X. Wang, Y. D. Li, Angew. Chem. 2003, 115, 3621; Angew. Chem. Int. Ed. 2003, 42, 3497.

- [12] K. N. Tu, U. Gosele, Appl. Phys. Lett. 2005, 86, 09311.
- [13] V. F. Puntes, K. M. Krishnan, A. P. Alivisatos, Science 2001, 291, 2115
- [14] H. van der Heide, R. Hemmel, C. F. van Bruggen, C. Haas, J. Solid State Chem. 1980, 33, 17.
- [15] C. E. M. Campos, J. C. de Lima, T. A. Grandi, K. D. Machado, V. Drago, P. S. Pizani, Solid State Commun. 2004, 131, 265.
- [16] B. Gates, Y. Yin, Y. Xia, J. Am. Chem. Soc. 2000, 122, 12582.
- [17] K.-S. Cho, D. V. Talapin, W. Gaschler, C. B. Murray, J. Am. Chem. Soc. 2005, 127, 7140.
- [18] a) H. W. Gu, R. K. Zheng, X. X. Zhang, B. Xu, J. Am. Chem. Soc. 2004, 126, 5664; b) H. W. Gu, Z. M. Yang, J. H. Gao, C. K. Chang, B. Xu, J. Am. Chem. Soc. 2005, 127, 34.

1245